

Chemistry Chapter 6 Test

Explain why the neutralization reaction of a weak acid and a strong base gives a weakly basic solution

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Explain why the ionization constant, K , for H_2SO_4 is larger than the ionization constant for HSO_4^-

Spherical Videos

Intro

The hydronium ion concentration in a sample of rainwater is found to be $1.7 \times 10^{-4} \text{ M}$ at 25°C . What is the concentration of hydroxide ion in the rainwater?

Is the self-ionization of water endothermic or exothermic? The ionization constant for water (K) is 2.9×10^{-14} at 40°C and 9.3×10^{-14} at 60°C .

What is the conjugate acid of each of the following? What is the conjugate base of each?

WORKS CITED

HOW TO CLASSIFY

Write equations that show NH_3 as both a conjugate acid and a conjugate base.

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Explain why the neutralization reaction of a weak acid and a strong base gives a weakly basic solution

General

What are amphiprotic species? Illustrate with suitable equations.

Explain why the neutralization reaction of a strong acid and a weak base gives a weakly acidic solution.

REVIEW Q'S

VSEPR • electron pairs have repulsion

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METALLIC BONDING

TYPES OF CHEMICAL

Chapter 6 Review - Chapter 6 Review 20 minutes - Covalent Bonding and Metallic Bonding.

OCTET RULE, E-DOT NOTATION, LEWIS STRUCTURES

Explain why the ionization constant, K , for HI is larger than the ionization constant for HF .

Calculate the pH and the pOH of each of the following solutions at 25°C which the substances ionize completely

Intro

Subtitles and closed captions

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State which of the following species are amphiprotic and chemical equations illustrating the amphiprotic character of

Chapter 6 Stoichiometry | MCQs | Class 9 Chemistry Federal Board New Book 2025 NBF - Chapter 6 Stoichiometry | MCQs | Class 9 Chemistry Federal Board New Book 2025 NBF 14 minutes, 7 seconds - MCQs | Class 9 **Chemistry**, Unit 06 Stoichiometry | **Chapter**, 06 Stoichiometry | Class 9 **Chemistry**, Federal Board 2025 New Book.

(SEC 5) INTERMOLECULAR FORCES

What are the hydronium and hydroxide ion concentrations in a solution whose pH is 6.52?

COVALENT BONDS

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